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multiplying the atomic ratios and atomic masses. You can see that a sample of N_2O weighing 44.02 grams contains 28.02 g of nitrogen and 16.00 g of oxygen. The mass percent of each element is calculated from its relative mass divided by the sum of the relative masses. Chemical compounds with integral atomic ratios,

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like nitrous oxide, are described as stoichiometric compounds, and they permit many simple calculations.

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Vocabulary – Give definitions and
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Reactant / Reagent 3. Theoretical Yield
4. Actual Yield 5. Percent Yield 6. Excess
Reagent 7. Limiting Reagent 8. Law of
Conservation of Matter Short Answer –
explain the following: 9. What is conserved
in every balanced equation (multiple
answers)? 10.

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The best definition for stoichiometry is the simple one: it's a way to figure out how much stuff you're going to make in a chemical reaction, or how much stuff you'll need to make a chemical reaction do what you want. When we put it that way,

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stoichiometry isn't so bad. We can deal with the crazy name if it's that simple.

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From aluminum to xenon, we explain the properties and composition of the substances that make up all matter.

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Answer Key

In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false.

_____true_____ 1. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical

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reaction is called stoichiometry.

_____true_____ 2.

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and products in a chemical reaction.

Different units can be used in these ratios, such as mass (grams,...

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Stoichiometry Questions. 1. Copper has two stable isotopes $^{63}_{29}\text{Cu}$ (atomic mass

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= 62.93 amu; 69.09% natural abundance)
and $^{65}_{29}\text{Cu}$ (atomic mass = 64.9278
amu; 30.91 % natural abundance).

Calculate the average atomic mass of
copper. 2. How many atoms of sulfur are
in 25.1 g of S? 3. Calculate the molecular
mass (in amu) of caffeine ($\text{C}_8\text{H}_{10}\text{N}_4\text{O}_2$). 4.

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Chemical Reactions According to the
College Board, "This unit explores
chemical transformations of matter by
building on the physical transformations
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involve the making and
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