

## Holt Chemistry Answer Key Acids And Bases

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An acid is a compound that donates a proton. After the proton leaves, the remaining compound (with one less proton) is the conjugate base. Example:  $\text{NH}_4^+ + \text{H}_2\text{O} \rightarrow \text{NH}_3 + \text{H}_3\text{O}^+$  Acid =  $\text{NH}_4^+$  Conjugate base =  $\text{NH}_3$

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reactants is the Brønsted-Lowry acid, and which is the Brønsted-Lowry base? Explain your answer.  $\text{H}_2\text{CO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \leftrightarrow \text{HCO}_3^-(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$ . The  $\text{H}_2\text{CO}_3$  is the Brønsted-Lowry acid because it donates a proton to the  $\text{H}_2\text{O}$ . The  $\text{H}_2\text{O}$  is the Brønsted-Lowry base because it accepts the proton. b.

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16. nitrous acid 17. hydrochloric acid 18. carbonic acid 19. sulfuric acid 20. hydriodic acid 21. hypobromous acid 22.  $\text{H}_2\text{S}$  23.  $\text{HNO}_3$  24.  $\text{H}_3\text{PO}_3$  25.  $\text{HClO}_4$  26.  $\text{NH}_3$  4 27. acidic:  $\text{HCl}$  and  $\text{NH}_4^+$  basic:  $\text{NH}_3$  and  $\text{Cl}^-$  28.  $\text{H}_2\text{O}$  and  $\text{OH}^-$   $\text{NH}_3$  and  $\text{NH}_4^+$  29. proton donors:  $\text{H}_2\text{O}$  and  $\text{NH}_4^+$  proton acceptors:  $\text{OH}^-$  and  $\text{NH}_3$  30.  $\text{HCl}(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow \text{NaCl}(\text{aq}) + \text{H}_2\text{O}(\text{l})$  31.  $\text{H}_3\text{O}^+(\text{aq}) + \text{Cl}^-(\text{aq}) + \text{Na}^+(\text{aq})$

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