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Equations
Answer Key

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Tutorial

Implications of a
Balanced Chemical
equation *Balancing*

Equations With
Polyatomic Ions

Anyone can do

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Balancing complex
chemical equations

How to Balance

Chemical Equations

How to Balance

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Method Writing
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Writing Skeleton
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Equilibrium: Crash
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#28 *Balancing
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~~Balancing Equations~~
Advanced: *Balancing
Equations Practice*

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Chemquest 29
Balancing Equations
Answer

1. First, let's balance the calcium atoms by placing a three in front of CaCl_2 : $3\text{CaCl}_2 + \text{Na}_3\text{N} \rightarrow 3\text{CaCl}_2 + \text{Na}_3\text{N}$

2. Next, let's balance the nitrogen atoms, by placing a 2 in front of Na_3N : $3\text{CaCl}_2 + 2\text{Na}_3\text{N} \rightarrow 3\text{CaCl}_2 + 2\text{Na}_3\text{N}$

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$3 \text{N}_2 + \text{NaCl} \rightarrow 3 \text{CaCl}_2 + 2 \text{Na}_3\text{N}$. 3. Now we will balance the sodium atoms by placing a 6 in front of NaCl.

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Name:

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Balancing Equations

Answer Key

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Balancing equations

answer. 1. First, let's

balance the calcium

atoms by placing a

three in front of CaCl

2 : Ca 3 N 2 + NaCl 3

CaCl 2 + Na 3 N. 2.

Next, let's balance

the nitrogen atoms, by

placing a 2 in front of

Na 3 N: Ca 3 N 2 +

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Equations Answers |

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3 a) Al₂O₃: 3 u): Nuro
lte c) 4 Na₂SO₄, 3C al
5 Mg₃(O)₂.

Information: How To
Balance Equations.

Consider the reaction
of sodium and
chlorine to produce
sodium chloride from
question one:

$\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$.

Remember that when

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chlorine is by itself it
is always written as
Cl₂ .
Answer Key

SslgDcII?g Ea1ue -

LTHS Answers

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Eventually, you will no

question discover a

extra experience and

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that you require to get
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similar to having
significantly cash?
Why don't you try to
acquire something
basic in the
beginning?

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the same exponents.
Your answer will them
have the same
exponent as the

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numbers you add or subtract. Here are some examples.

Make sure you understand each step:

$$4.2 \times 10^6 + 3.1 \times 10^5$$

make exponents the same, either a 5 or 6

$$42 \times 10^5 + 3.1 \times 10^5 =$$
$$45.1 \times 10^5 = 4.51 \times 10^6$$

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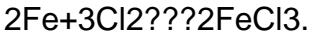
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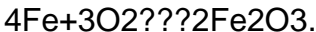
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1. Balanced
equations.

(Coefficients equal to
one (1) do not need to
be shown in your
answers). (a)



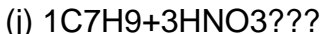
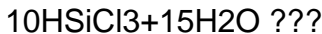
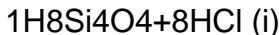
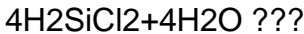
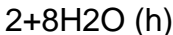
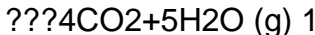
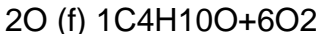
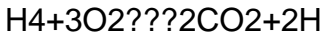
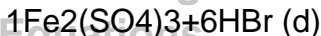
(b)



(c)

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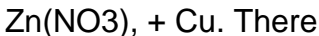
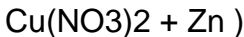
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Answer Key

Balancing Equations:

Practice Problems



There is a sentence that you

can write to describe

what happens above:

Copper(II) nitrate

reacts with zinc metal

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and yields zinc nitrate and copper metal.

Take a minute to notice the relationship between the sentence and the symbols in the reaction. Now, try a few yourself.

Intro, to chemical
Reoctioil\$ - LTHS
Answers

Hint: add your answer

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to question 1 and four times your answer to question 2. $4 + 4(7) = 4 + 28 = 32$. 4. Verify

that there are 32 electrons pictured in the drawing of CCl_4 . Count each dot in the picture above: there are 24. There are also 8 more electrons represented by the 4 lines (each line represents two

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Balancing
Equations

ChemQuest 23:

Drawing Covalent
Compounds

balancing equations
answers key. The first
step to balance the
equation is to write
down the chemical
formula of reactants
that are listed on the
left side of the

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chemical equation.

After this, you can list down the products on the right hand side of the chemical equation.

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