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|•74 L On15L 0-59L Ending Buret Reading: 11.83 L10-10L  
10.591 mL of Titrant Used: 10.09L/9.95L10.00L mmole of  
NaOH: 1.009 mol 9.95 no 1.0 ml mmole of H<sub>2</sub>SO<sub>4</sub> 0.5 mal  
5.0 0.5 5.0nd Molarity of H<sub>2</sub>SO<sub>4</sub> Solution 0.05 M 0.5 M  
0.05 M 0.05M Average Molarity 0.6M.

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Reaction 1:  $1 \text{ NaHCO}_3 (\text{s}) + 1 \text{ NaOH} (\text{s}) + 1 \text{ CO}_2 (\text{g})$  Reaction

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2:  $2 \text{NaHCO}_3 (\text{s}) \rightarrow 1 \text{Na}_2\text{CO}_3 (\text{s}) + 1 \text{CO}_2 (\text{g}) + 1 \text{H}_2\text{O} (\text{g})$

Reaction 3:  $2 \text{NaHCO}_3 (\text{s}) \rightarrow 1 \text{Na}_2\text{O} (\text{s}) + 2 \text{CO}_2 (\text{g}) + \text{H}_2\text{O} (\text{g})$ . Propose a voltaic cell that would coat a piece of copper with platinum. Include the half reactions for your cell and the voltage potential based on the.

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1. Identify the group for each unknown. Your unknowns will be from three different groups. (Example: A  $K^+$  salt, a halogen, and an acid.) 2. Identify the specific compound from each group. (Example:  $KCl$ ,  $Br_2$  and  $HCl$ ) 3. Explain the reasoning that enabled you to come to these conclusions.

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