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physical Page 6/9 Chapter 18 Chemical Equilibrium Worksheet
Answers CHAPTER 18 Chemical Equilibrium At equilibrium, the
rate of the forward reaction is equal to the rate of the reverse
reaction. The equilibrium constant, K , is used to determine the
relative concentrations of products and reactants at equilibrium.
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1. $A + B \rightleftharpoons C + D$ (forward reaction) $C + D \rightleftharpoons A + B$ (reverse reaction)
Equilibrium (forward rate = reverse rate) remain constant. The ratio of the mathematical product $[C]^x[D]^y$ to the mathematical product $[A]^n[B]^m$ for this reaction has a definite value at a given temperature.

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Chapter 18 of how fast a reaction occurs. 14: Rates of Chemical Reactions - Chemistry LibreTexts As before, there are three

reaction rates in this reaction: k_1 , k_{-1} , and k_2 . The pre-equilibrium

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approximation uses the rate constants to solve for the rate of the reaction, indicating how

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Chapter 18 Chemical Equilibrium Worksheet sometimes referred to as the chemical-equilibrium expression. The H_2 , I_2 , HI

Equilibrium System Consider the reaction between H_2 and I_2 vapor in a sealed flask at an elevated temperature. The rate of reaction can be followed by observing the rate at which the violet color of the iodine vapor diminishes, as $[C]^x[D]^y [A]^n[B]^m$ 556

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DorettaA TEACHER. Modern Chemistry: Chapter 18. Chemical Equilibrium. Decreases. Increases. Equilibrium Constant Expression. state reached by a chemical reaction where there is no change. As equilibrium is approached, the rate of the forward reaction. As equilibrium is approached, the rate of the reverse reaction.

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Chapter 18.1 : The Nature of Chemical Equilibrium 1. Chapter 18.1
2. 1. Define chemical equilibrium. 2. Explain the nature of the equilibrium constant. 3. Write chemical equilibrium expressions and carry out calculations involving them. 3. □ Theoretically, every reaction can proceed in two directions, forward and reverse.

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Chapter 18 Notes Reaction Rates and Equilibrium. 18.1 Rates of Reaction. Collision Theory o Rate = The speed of any change that occurs within an interval of time o KEY = In chemistry, the rate of chemical change or the reaction rate is usually expressed as the

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amount of reactant changing per unit time o Collision Theory = atoms,

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