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Solubility product =  $K_{sp}$  Consider the equilibrium between undissolved silver chloride and a saturated silver chloride solution:  $\text{AgCl}(s) \rightleftharpoons \text{Ag}^+(aq) + \text{Cl}^-(aq)$

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Since this is a heterogeneous equilibrium, the equilibrium constant is  $K_{sp} = [Ag^+][Cl^-]$  always write the  $K_{sp}$  expression with the solid on the left

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At a certain temperature, the solubility of  $Fe(OH)_2$  in water is  $7.7 \times 10^{-6}$  mol/L (M). Its  $K_{sp}$  can be calculated based on the equilibrium equation:

$Fe(OH)_2 \rightleftharpoons Fe^{2+} + 2OH^-$   
 $Fe(OH)_2 \rightleftharpoons Fe^{2+} + 2OH^-$   
Therefore, the solubility product expression is:  $K_{sp} = [Fe^{2+}][OH^-]^2$   
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