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Atomic Structure: Protons, Electrons

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A subatomic particle with a negative  
charge. electron. The smallest known

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particles that make up protons and neutrons. quarks. The positively charged center of that atom where most of the mass of the atom is located. nucleus. The negatively charged area around the nucleus where the volume of the atom is located.

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## STUDY GUIDE ANSWERS ATOMIC STRUCTURE - PERIODIC TABLE

TEST 1. How many periods make up the periodic table? 7 2. How many groups or families make up the periodic table? 9 families (The Lanthanide Series and Actinide Series are part of the transition metals family.) There are 18 groups.

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GUIDE ANSWERS ATOMIC ...

Name \_\_\_\_\_

\_\_\_\_\_ Block \_\_\_\_\_ Test

Date \_\_\_\_\_ . Atomic Structures

Study Guide ANSWERS. 1. The Periodic

Table is mostly made of metals. 2. The

noble gases are the farthest group to the

right on the Periodic Table. 3. Atoms

having equal numbers of positive and

negative charges make them neutral.

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Atomic Structures Study Guide

ANSWERS

Calculate the average atomic mass for Ne

given the following isotopes & their

abundance: Neon-20 = 90.92%; Neon-21

= 0.257%; Neon-22 = 8.82%. Study Guide

Answer Key. Answer the following

questions regarding the discovery of the

atom & its parts. A. Identify 3 conclusions

of the gold foil experiment.

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Atomic Structures Study Guide

ANSWERS 1. The Periodic Table is mostly made of metals. 2. The noble

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gases are the farthest group to the right on the Periodic Table. 3. Atoms having equal numbers of positive and negative charges make them neutral. 4.

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## Atomic Structure Study Guide Answers

1. Describe the evolution of atomic theory leading to the current model of the atom based on the works of Democritus, Dalton, Thomson, Rutherford, Bohr, and Schrodinger. 2. Explain the law of conservation of mass, the law of definite proportions, and the law of multiple proportions. 3. Describe the properties of protons, neutrons, and electrons.

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## Atomic Structure And The Periodic Table Study Guide ...

CH1010 Exam #2 Study Guide For reference see □Chemistry: An Atoms-



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focused Approach by Gilbert, Kirss, and Foster Chapter 3: Atomic Structure, Explaining the Properties of Elements Trends to know (and be able to explain the trend—think about  $Z_{\text{eff}}$ ): o Effective Nuclear Charge ( $Z_{\text{eff}}$ ): the attraction toward the nucleus experienced by

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Chapter 3: Atomic Structure, Explaining the Properties of ...

Atoms are the smallest unit of matter that cannot be divided using any chemical method. They do consist of smaller parts, but can only be broken by nuclear reactions. The three parts of an atom are protons, neutrons, and electrons. Protons carry a positive electrical charge. Neutrons are electrically neutral.

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Basic Atomic Structure and Atomic

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Theory - Study Guide

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Atomic Structure Study Guide Answers  
ATOMIC STRUCTURE: SUBATOMIC  
PARTICLE CHARGE LOCATION  
WITHIN THE ATOM ATOMIC MASS  
(AMU) Proton 15. 16. 17.

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 $x = \text{fraction } ^{37}\text{Cl}$   $1 - x = \text{fraction } ^{35}\text{Cl}$ .  
Instead of using the integers 37 and 35 as atomic masses, take the more precise atomic masses of the isotopes from Table 1: The calculation reveals that natural chlorine is 24%  $^{37}\text{Cl}$  and 76%  $^{35}\text{Cl}$ .

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